

Chemistry 210 – General Chemistry II

Winter/Spring 2008

11:30-12:20 MWF, SL118

Dr. Jeffrey J. Bodwin

bodwin@mnstate.edu

407H Hagen Hall

477-4371 (office)

<http://www.mnstate.edu/bodwin/>

Office hours:

Office hours subject to change, check web page

Required Material:

“Chemistry: A Molecular Approach”; Tro
Mastering Chemistry {MC} System access (online homework system)

CHEM 210 General Chemistry II [B1] (4)

General chemistry principles: kinetics, chemical equilibrium, acid-base chemistry, solubility equilibrium, thermodynamics, oxidation-reduction, electrochemistry, coordination chemistry, and nuclear chemistry. Lab included. Prerequisite: CHEM 150

Class E-Mail List:

chem210jb@mnstate.edu

An email listserv has been created for this class. It will be used for class announcements including Mastering General Chemistry assignment. It is also where I will respond to all email questions. If you have questions, you can either email them to the list or to me directly. All relevant questions mailed to me directly will be answered to the list with the questioner's identity removed. To subscribe to the list, send an email to “majordomo@mnstate.edu” with “subscribe chem210jb” in the body. This should be done as soon as practical from the email account you are most likely to check on a regular basis.

Grading:

Grades will be based upon 3 of 4 exams (150pts each, tentative dates Jan. 28, Feb. 22, Mar. 21, Apr. 25), MC assignments (100pts), a final exam (200pts), and the lab grade (250pts).

Exams	3 x 150 = 450pts
MC	100pts
Final Exam	200pts
<u>Lab grade</u>	<u>250pts</u>
<i>Total Points</i>	<i>1000pts</i>

Tentative grade assignments are: A = 90-100%, B = 80-90%, C = 70-80%, D = 60-70%. These cutoffs *may* be lowered at the instructor's discretion, but they will not be raised.

You lab grade is 25% of your course grade and is determined by your lab instructor. At the end of the semester, your lab and lecture grades will be combined using the point breakdown shown above, with your performance in lab scaled to 250 points. If you do not receive a passing grade in the lab portion of the class, you will receive a grade of “F” for the course regardless of exam performance.

Regular and punctual attendance is expected and may be recorded. Late arrival on exam days is not acceptable as it disturbs those who arrive on time; therefore, no exams will be distributed after the test period has begun. If you anticipate that this will be a problem, let me know **BEFORE** the exam. There will be no make-up exams. Exams will be closed book and a **non-graphing** calculator will typically be allowed. The Final Exam will be cumulative. If you do not

take the final exam, you will receive a grade of “F” for the course regardless of previous performance.

Mastering Chemistry assignments will be made on a regular basis relating to the current lecture material. Sufficient time will be given between lecture and the MC deadlines that no extensions will be required. Although MC is a very useful tool, it is not sufficient to *only* do the MC problems. Regularly attempting the problems in the text will also be required for your success.

Academic Honesty

Cheating will not be tolerated and will be reported to the Dean of your College and the Vice President for Academic Affairs. It may also be reported to the Student Conduct Committee for further disciplinary action. For a full description of the MSUM Academic Honesty Policy, please see the Student Handbook. {<http://www.mnstate.edu/sthandbook/POLICY/index.htm>}

Disability Access Statement: Students with disabilities who believe they may need an accommodation in this class are encouraged to contact Greg Toutges, Coordinator of Disability Services at 477-5859 (Voice) or 1-800-627-3529 (MRS/TTY), CMU 114 as soon as possible to ensure that accommodations are implemented in a timely fashion.

Tentative Lecture Schedule

Dates	Chapter
Jan. 9-11	Ch. 5 – Gases
Jan. 14-18	Ch. 11 – Liquids, Solids and Intermolecular Forces
Jan. 18-25	Ch. 12 – Solutions
Jan. 28	Exam 1
Jan. 30-Feb. 8	Ch. 13 – Chemical Kinetics
Feb. 11-20	Ch. 14 – Chemical Equilibrium
Feb. 22	Exam 2
Feb. 25-29	Ch. 15 – Acids and Bases
Mar. 10-19	Ch. 16 – Aqueous Ionic Equilibrium
Mar. 21	Exam 3
Mar. 26-Apr. 7	Ch. 17 – Free Energy and Thermodynamics
Apr. 9-18	Ch. 18 – Electrochemistry
Apr. 21-23	Ch. 19 – Radioactivity and Nuclear Chemistry
Apr. 25	Exam 4
Apr. 28	Review
May 5	Final Exam, noon