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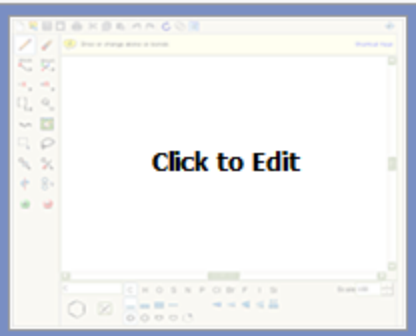
- **Chemical Structure**
- Markush
- Molecular Formula
- Property NEW
- Substance Identifier

Chemical Structure



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Characteristic(s)

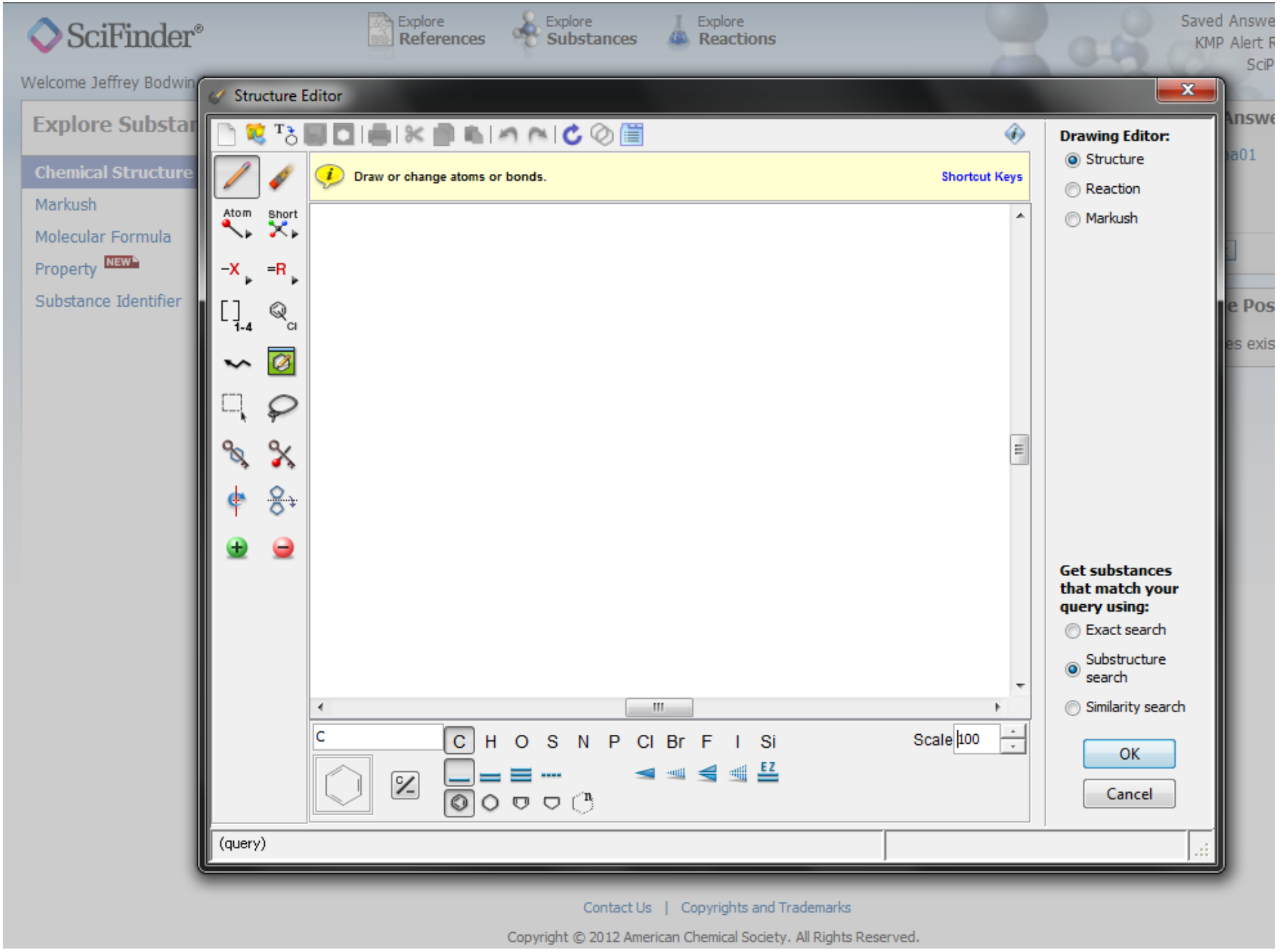
- ☐ Single component
- ☐ Commercially available
- ☐ Included in reference(s)

Class(es)

- ☐ Alloys
- ☐ Coordination compounds
- ☐ Incompletely defined
- ☐ Mixtures
- ☐ Polymers
- ☐ Organics, and others not listed

Studies

- ☐ Analytical
- ☐ Biological
- ☐ Preparation
- ☐ Reactant or reagent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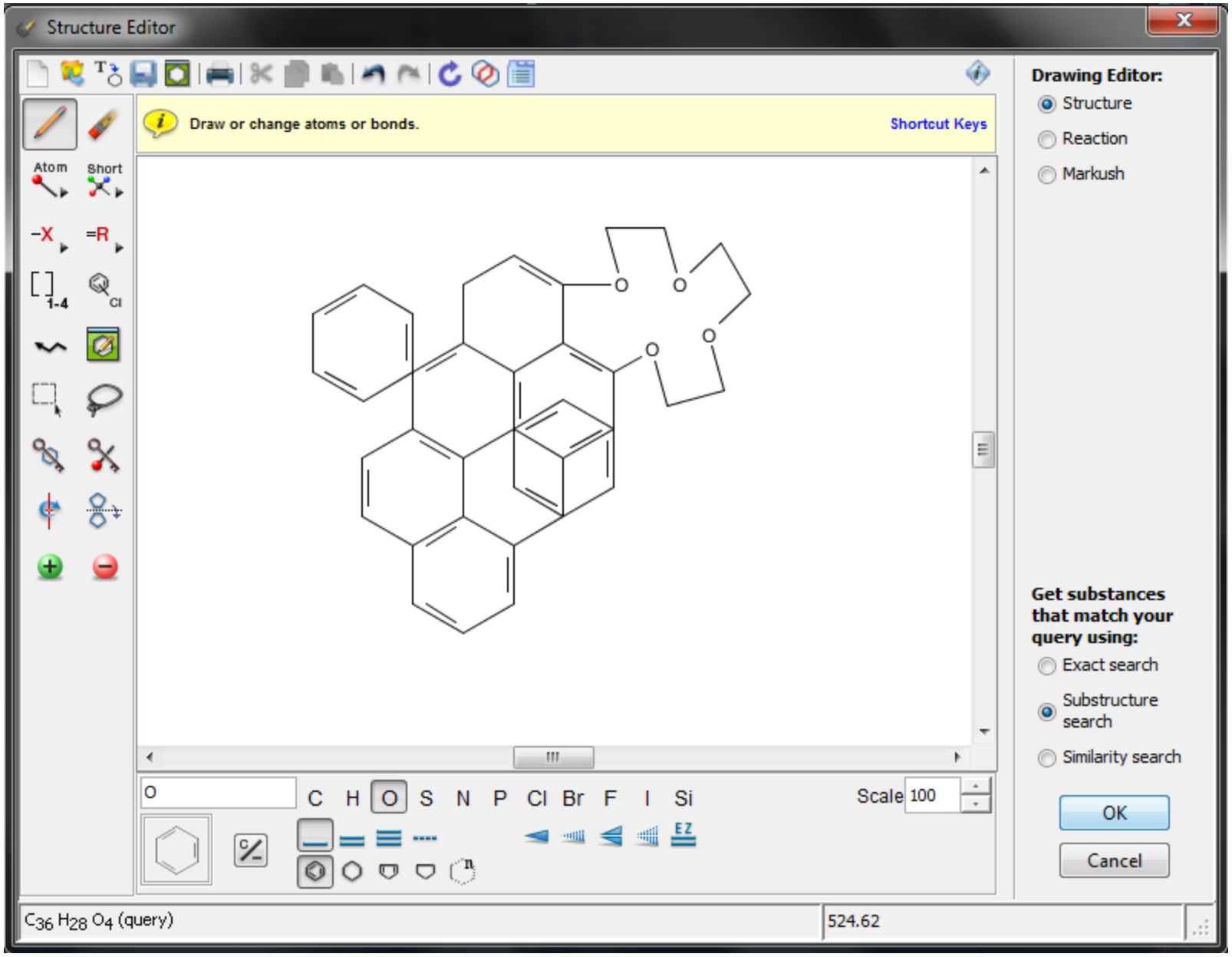
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Similarity search
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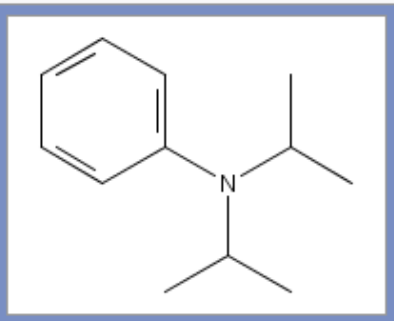
Chemical Structure

- Markush
- Molecular Formula
- Property NEW
- Substance Identifier

Chemical Structure



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Search type:
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- Similarity

Show precision analysis

Characteristic(s)

- Single component
- Commercially available
- Included in reference(s)

Class(es)

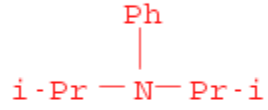
- Alloys
- Coordination compounds
- Incompletely defined
- Mixtures
- Polymers
- Organics, and others not listed

Studies

- Analytical
- Biological
- Preparation
- Reactant or reagent

1. **Substance Detail**
4107-98-6

~89

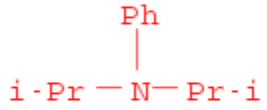


$C_{12}H_{19}N$
Benzenamine, *N,N*-bis(1-methylethyl)-

Experimental Properties

2. **Substance Detail**
106834-37-1
(Component: 4107-98-6)

~9

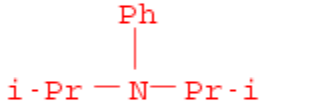


$C_{12}H_{19}N \cdot H$
Benzenamine, *N,N*-bis(1-methylethyl)-, conjugate acid (1:1)

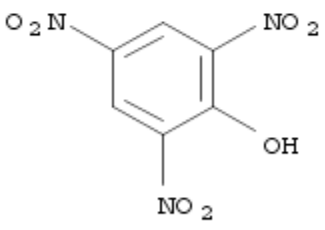
3. **Substance Detail**
23970-60-7

~3

4107-98-6
$C_{12}H_{19}N$



88-89-1
$C_6H_3N_3O_7$



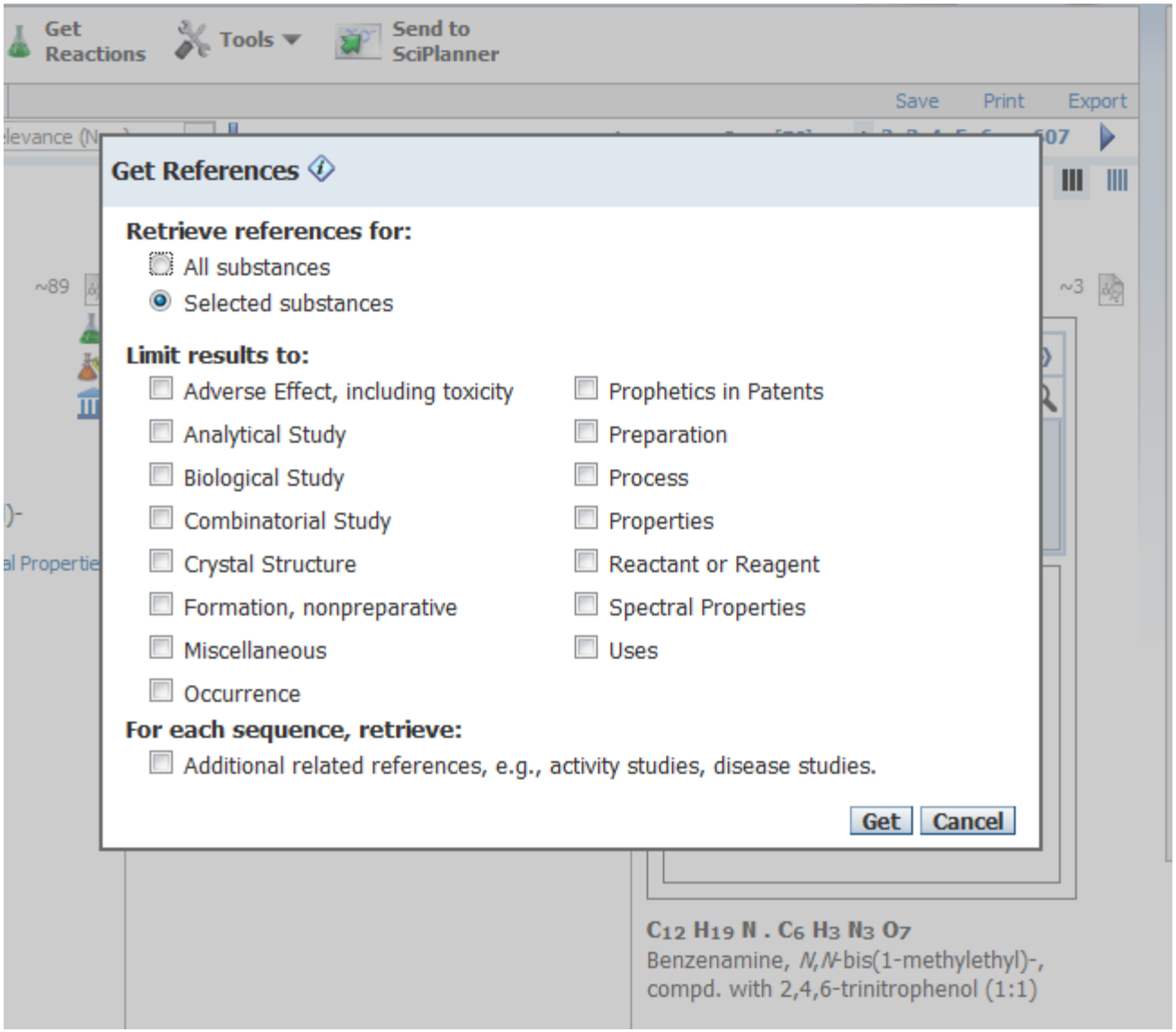
$C_{12}H_{19}N \cdot C_6H_3N_3O_7$
Benzenamine, *N,N*-bis(1-methylethyl)-, compd. with 2,4,6-trinitrophenol (1:1)

Experimental Properties

4. **Substance Detail**

5. **Substance Detail**

6. **Substance Detail**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C12 H19 N . C6 H3 N3 O7
Benzenamine, N,N-bis(1-methylethyl)-, compd. with 2,4,6-trinitrophenol (1:1)

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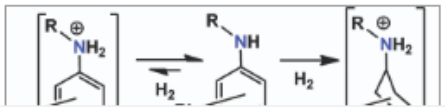
Display:

1. **Preparation method of n-alkylarylamine** Full Text
By Yan, Xinhuan; Yang, Fang
From Faming Zhuanli Shenqing (2012), CN 102531912 A 20120704. | Language: Chinese, Database: CAPLUS
~0
The title method comprises mixing arom. nitro compd. with C1-C6 fatty alc. and water, reacting in the presence of PtSn/Al2O3 catalyst at 180°C and 2 MPa in fixed-bed reactor, rotatory evapg., adsorbing on silica gel column and eluting with Et ether-Et acetate to obtain N-alkylarylamine. The arom. nitro compd. is nitrobenzene, o-methylnitrobenzene, p-methylnitrobenzene, o-methoxynitrobenzene, p-methoxynitrobenzene, etc. The PtSn/Al2O3 catalyst is prepd. by (1) calcining γ-Al2O3 at 500°C for 3 h to obtain support, soaking in SnCl2·H2O/water soln., stirring, evapg., calcining to obtain Sn/γ-Al2...

2. **Anti-slip composition using dispersing agent for pavement of road** Full Text
By Choi, Il Gwang; Kil, Min Ho; Jeon, Eun Jin
From Repub. Korea (2012), KR 1132571 B1 20120405. | Language: Korean, Database: CAPLUS
~0
The title compn. comprises 100 wt. parts of mixt. contg., by wt.%, Me methacrylate 30-50, (meth)acrylate monomer except Me methacrylate 30-40, and acrylic polymer 20-40, 0.1-2 wt. parts of dispersing agent, 0.1-5 wt. parts of wax, 0.1-5 wt. parts of crosslinking agent, 1-10 wt. parts of polymn. promoter, 0.005-0.1 wt. part of polymn. preventing agent, and 1-20 wt. parts of polymn. initiator. The compn. has ensured reaction stability in summer, good mech. properties such as tensile strength and elongation percentage, and good volatile inhibition ability. Due to good volatile inhibition abilit...

3. **Synthesis of Hindered Anilines: Copper-Catalyzed Electrophilic Amination of Aryl Boronic Esters** Full Text
By Rucker, Richard P.; Whittaker, Aaron M.; Dang, Hester; Lalic, Gojko
From Angewandte Chemie, International Edition (2012), 51(16), 3953-3956, S3953/1-S3953/108. | Language: English, Database: CAPLUS
~2
The authors have developed a mild copper-catalyzed electrophilic amination reaction for the synthesis of sterically-hindered anilines from aryl and heteroaryl boronic esters. The new method is compatible with a wide range of functionalities, including chloro, bromo, iodo, carbomethoxy, nitro, hydroxyl, formyl, and methoxy groups. Overall, an exceptionally broad scope and reliability of this new procedure, together with the availability of a wide variety of aryl boronic esters, make it a significant addn. to the existing methods for aniline synthesis.

4. **Metal-Free Aromatic Hydrogenation: Aniline to Cyclohexyl-amine Derivatives** Full Text


By Mahdi, Tayseer; Heiden, Zachariah M.; Grimme, Stefan; Stephan, Douglas W.
From Journal of the American Chemical Society (2012), 134(9), 4088-4091.
| Language: English, Database: CAPLUS
~4
Hydrogenation of the N-bound Ph rings of amines, imines, and aziridine is achieved

Analysis | Refine

Refine by:

- Research Topic
- Author Name
- Company Name
- Document Type
- Publication Year
- Language
- Database

Research Topic

Examples:

The effect of antibiotic residues on dairy products

Photocyanation of aromatic compounds

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5. Studies on the atmospheric fate of propachlor (2-chloro-N-isopropylacetanilide) in the gas-phase

By: Munoz, Amalia; Vera, Teresa; Sidebottom, Howard; Rodenas, Milagros; Borras, Esther; Vazquez, Monica; Raro, Montserrat; Mellouki, Abdelwahid

The gas-phase degrdn. of propachlor (2-chloro-N-isopropylacetanilide), a widely used herbicide, was investigated under atm. conditions at the large outdoor European simulation chamber (EUPHORE) in Valencia, Spain. The rate coeff. for the reaction of hydroxyl radicals with propachlor was measured using a conventional relative rate technique. A value of the rate coeff. for the reaction of OH radicals with propachlor, k OH (propachlor) = (1.5 ± 0.3) × 10-11 cm3 mol.-1 s-1 was detd. at 300 ± 5 K and atm. pressure. Rate coeff. data for the reaction of OH radicals with N,N-diisopropylaniline and N-methylacetanilide, which have structural similarities to propachlor, were also obtained using the relative rate method, k OH(N,N-diisopropylaniline) = (4.4 ± 0.5) × 10-11 cm3 mol.-1 s-1 and k OH(N-methylacetanilde) = (2.7 ± 0.2) × 10-11 cm3 mol.-1 s-1, resp. The rate coeffs. for photolysis of propachlor, J(propachlor) < 2.0 × 10-5 s-1, and reaction of ozone with propachlor, kO3(propachlor) < 1.5 × 10-19 cm3 mol.-1 s-1, under atm. conditions were also detd. The available kinetic data suggest that the gas-phase tropospheric degrdn. of propachlor will be mainly controlled by reaction with OH and possibly NO3 radicals, and provide the basis of an est. for the propachlor tropospheric lifetime of approx. 20 h. The atm. implications of the use of this species as a herbicide are discussed.

Indexing

Air Pollution and Industrial Hygiene (Section59-2)

Section cross-reference(s): 53

Concepts

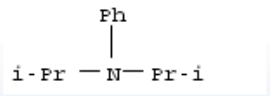
IR spectroscopy

Fourier-transform; studies on atm. fate of propachlor (2-chloro-N-isopropylacetanilide) in gas-phase

Air pollution
Atmospheric aerosols
Atmospheric chemistry
Herbicides

Substances

579-10-2 N-Methylacetanilide
1918-16-7 Propachlor
4107-98-6 N,N-Diisopropylaniline



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Pages33-40
Journal; Online Computer File
2012
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Company/Organization

Instituto Universitario UMH-CEAM
Paterna (Valencia), Spain 46980

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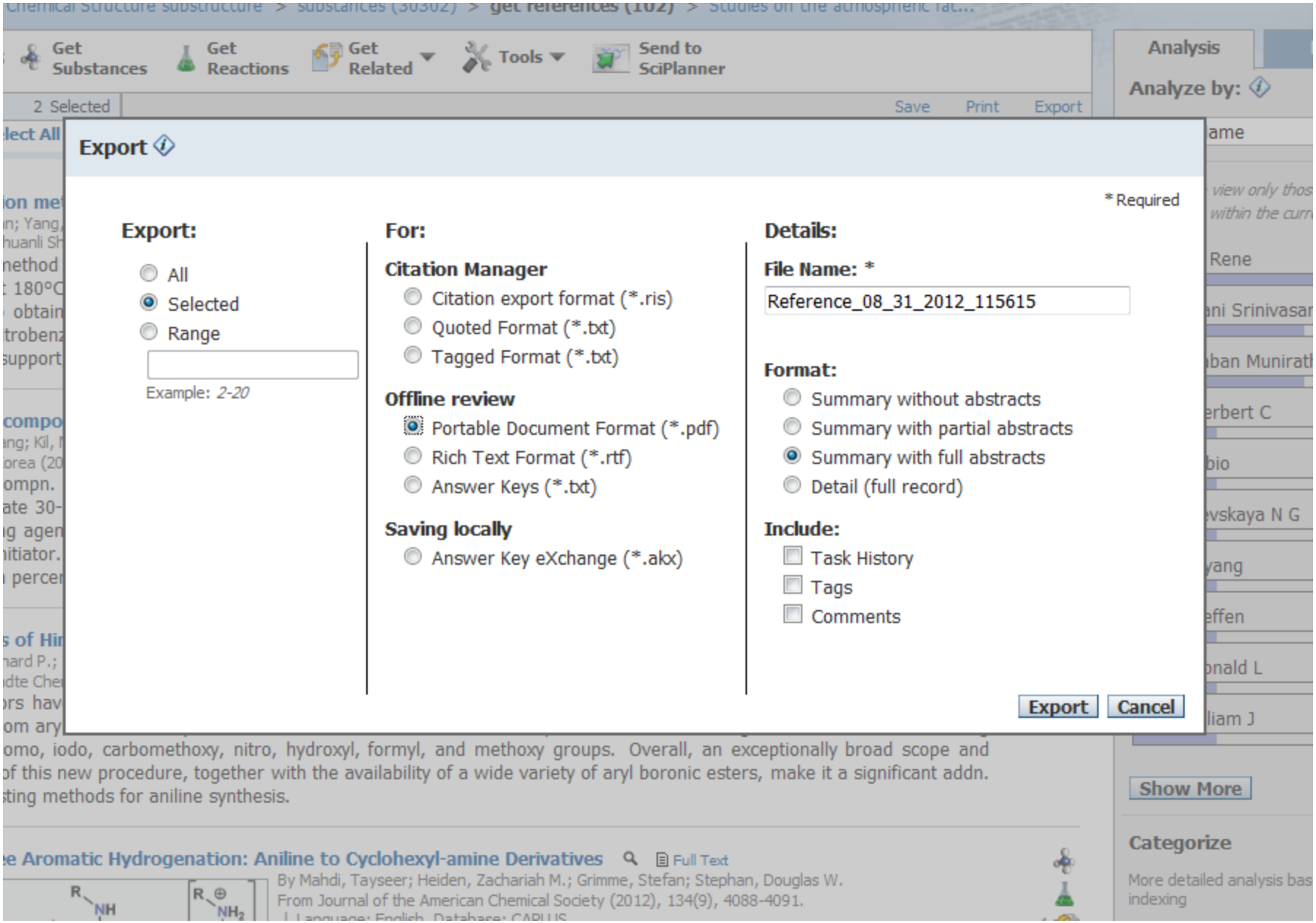
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By Mahdi, Tayseer; Heiden, Zachariah M.; Grimme, Stefan; Stephan, Douglas W.
From Journal of the American Chemical Society (2012), 134(9), 4088-4091.